

Series : SGRP1

SET~2

प्रश्न-पत्र कोड
Q.P. Code

56/1/2

रोल नं.

Roll No.

21602840

परीक्षार्थी प्रश्न-पत्र कोड को उत्तर-पुस्तिका के मुख-पृष्ठ पर अवश्य लिखें।

Candidates must write the Q.P. Code on the title page of the answer-book.

नोट	NOTE :
(I) कृपया जाँच कर लें कि इस प्रश्न-पत्र में मुद्रित पृष्ठ 23 हैं।	(I) Please check that this question paper contains 23 printed pages.
(II) प्रश्न-पत्र में दाहिने हाथ की ओर दिए गए प्रश्न-पत्र कोड को परीक्षार्थी उत्तर-पुस्तिका के मुख-पृष्ठ पर लिखें।	(II) Q.P. Code given on the right hand side of the question paper should be written on the title page of the answer-book by the candidate.
(III) कृपया जाँच कर लें कि इस प्रश्न-पत्र में 33 प्रश्न हैं।	(III) Please check that this question paper contains 33 questions.
(IV) कृपया प्रश्न का उत्तर लिखना शुरू करने से पहले, उत्तर-पुस्तिका में यथा स्थान पर प्रश्न का क्रमांक अवश्य लिखें। {}	(IV) Please write down the serial number of the question in the answer-book at the given place before attempting it.
(V) इस प्रश्न-पत्र को पढ़ने के लिए 15 मिनट का समय दिया गया है। प्रश्न-पत्र का वितरण पूर्वाह्न में 10.15 बजे किया जाएगा। 10.15 बजे से 10.30 बजे तक परीक्षार्थी केवल प्रश्न-पत्र को पढ़ेंगे और इस अवधि के दौरान वे उत्तर-पुस्तिका पर कोई उत्तर नहीं लिखेंगे।	(V) 15 minute time has been allotted to read this question paper. The question paper will be distributed at 10.15 a.m. From 10.15 a.m. to 10.30 a.m., the candidates will read the question paper only and will not write any answer on the answer-book during this period.



रसायन विज्ञान (सैद्धांतिक)
CHEMISTRY (Theory)



निर्धारित समय : 3 घण्टे

Time allowed : 3 hours

अधिकतम अंक : 70

Maximum Marks : 70

^56/1/2^

2309-2

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[P.T.O.]

$$k = \frac{2.303}{t} \log \frac{A_0}{A}$$

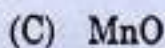
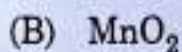
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SECTION - A

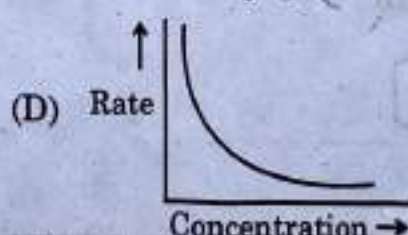
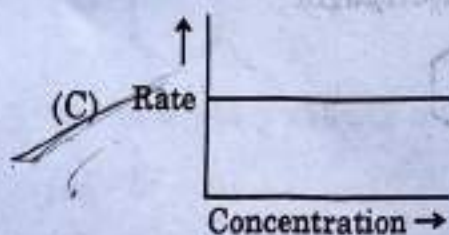
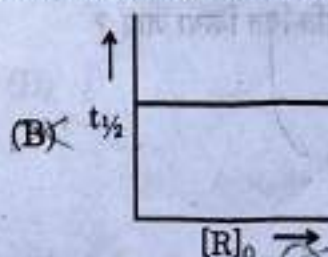
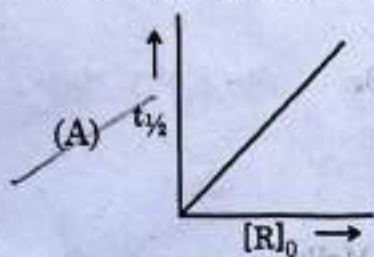
$$\log \frac{A_0}{A} = \frac{2.303}{t} \times \text{Area under curve}$$

Questions No. 1 to 16 are Multiple Choice type questions carrying 1 mark each.

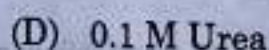
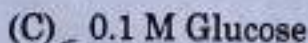
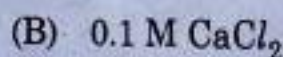
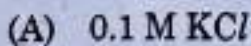
1. Which of the following is most basic?



2. Which of the following curve represents the first order reaction?



3. Which of the following solutions will have the highest osmotic pressure?



4. For a reaction : $\text{N}_2 + 3\text{H}_2 \longrightarrow 2\text{NH}_3$, the rate of reaction with respect to NH_3 is

(A) $+\frac{1}{3} \frac{\Delta [\text{NH}_3]}{\Delta t}$

(B) $-\frac{1}{2} \frac{\Delta [\text{NH}_3]}{\Delta t}$

(C) $+\frac{1}{4} \frac{\Delta [\text{NH}_3]}{\Delta t}$

(D) $+\frac{1}{2} \frac{\Delta [\text{NH}_3]}{\Delta t}$



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5. How many Faradays are required to reduce 1 mol of $\text{Cr}_2\text{O}_7^{2-}$ to Cr^{3+} in acidic medium ?

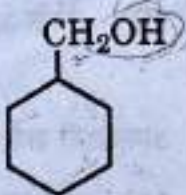
(A) 2

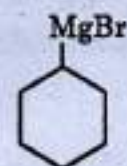
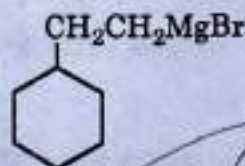
(B) 3

(C) 6

(D) 4

1

6. Which of the following Grignard Reagent will be used to prepare  when treated with methanal ?

(A) CH_3MgBr (B) (C) (D) 

1

7. The secondary valency of Pt in $[\text{Pt}(\text{en})_2\text{Cl}_2]^{2+}$ is

(A) 6

(B) 5

(C) 4

(D) 2

1

8. $(\text{CH}_3)_3\text{C} - \text{OC}_2\text{H}_5$ on reaction with HI gives

(A) $(\text{CH}_3)_3\text{C} - \text{I}$ and $\text{C}_2\text{H}_5 - \text{I}$ (B) $(\text{CH}_3)_3\text{C} - \text{OH}$ and $\text{C}_2\text{H}_5 - \text{I}$ (C) $(\text{CH}_3)_3\text{C} - \text{I}$ and $\text{C}_2\text{H}_5 - \text{OH}$ (D) $(\text{CH}_3)_3\text{C} - \text{OH}$ and $\text{C}_2\text{H}_5 - \text{OH}$

1

9. Aniline on direct nitration yields

(A) 51%-ortho, 47%-para, 2%-meta derivatives

(B) 51%-meta, 47%-ortho, 2%-para derivatives

(C) 51%-para, 47%-meta, 2%-ortho derivatives

(D) 51%-ortho, 47%-meta, 2%-para derivatives

1



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10. Which of the following is 'not' true about enantiomers ?

- (A) They have the same chemical reactivity.
(B) They have the same specific rotation.
(C) They have the same melting or boiling point.
(D) They have the same refractive index.

11. An example of non-reducing sugar is

- (A) Glucose (B) Sucrose
(C) Lactose (D) Maltose

12. Benzene diazonium chloride on reaction with phenol in weakly basic medium gives

- (A) Azobenzene (B) Benzene
(C) Chlorobenzene (D) p-hydroxyazobenzene

For questions number 13 to 16, two statements are given – one labelled as Assertion (A) and the other labelled as Reason (R). Select the correct answer to these questions from the codes (A), (B), (C) and (D) as given below :

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
(B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of the Assertion (A).
(C) Assertion (A) is true, but Reason (R) is false.
(D) Assertion (A) is false, but Reason (R) is true.

13. Assertion (A) : The presence of $-OH$ group in phenols directs the incoming group to ortho and para positions.Reason (R) : $-OH$ group in phenols activates the aromatic ring towards electrophilic substitution reaction.

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14. Assertion (A) : Actinoids show irregularities in their electronic configurations. 1

Reason (R) : In actinoids 5f, 6d and 7s orbitals are of comparable energies.

15. Assertion (A) : Components of azeotropes are easily separated by fractional distillation. 1

Reason (R) : Components of an azeotrope have same composition in liquid and vapour phase.

16. Assertion (A) : The two strands of DNA are complementary to each other. 1

Reason (R) : The hydrogen bonds are formed between specific pairs of bases.

SECTION - B

17. What type of deviation from Raoult's law is shown by mixture of ethanol and acetone ? Give reason. What will happen to the boiling point of the solution on mixing ethanol and acetone ? 2

18. (a) Write IUPAC names of the following coordination compounds : 2×1

(i) $[\text{Ag}(\text{NH}_3)_2][\text{Ag}(\text{CN})_2]$

(ii) $\text{K}_3[\text{Fe}(\text{C}_2\text{O}_4)_3]$

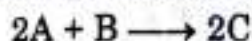
OR

(b) (i) Give a chemical test to show that $[\text{Co}(\text{NH}_3)_5\text{SO}_4]\text{Cl}$ and $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{SO}_4$ are ionisation isomers. 2×1

(ii) What is meant by the 'Chelate effect' ? Give an example.

19. Why are haloarenes less reactive towards nucleophilic substitution reaction ? Give two reasons. 2

20. Following reaction takes place in one step : 2



How will the rate of above reaction change if the volume of the reaction vessel is decreased to one third of its original volume ? Will there be any change in the order of reaction with the reduced volume ?



{

21. Differentiate between the following :

2 × 1

- (i) Acidic amino acids and basic amino acids
- (ii) Nucleotide and Nucleoside

SECTION - C

22. Compound 'X' with molecular formula C_4H_9Br reacts with aqueous KOH to give an alcohol. The rate of this reaction depends only on the concentration of the compound 'X'. When an optically active isomer 'Y' of the compound 'X' was treated with aqueous KOH solution, the rate of reaction was found to be dependent on concentration of compound 'Y' and aqueous KOH both.

3 × 1

- (a) Write down the structural formula of both 'X' and 'Y'.
- (b) Out of 'X' and 'Y', which one will undergo racemisation and why?
- (c) Out of 'X' and 'Y', which one will form product with inversion of configuration and why?

23. Write the reaction involved in the following :

3 × 1

- (a) Reimer-Tiemann reaction
- (b) Kolbe's reaction
- (c) Friedal-Crafts acylation of anisole

24. Give reasons for the following :

3 × 1

- (a) Carboxylic acids have higher boiling point than alcohols of comparable molecular masses.
- (b) Alpha (α) hydrogens of aldehydes and ketones are acidic in nature.
- (c) Nucleophilic addition of ammonia and its derivatives does not occur with carbonyl group in strongly acidic medium.



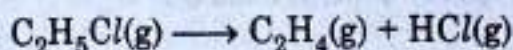
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25. Define the following terms :

- (a) Anomers
(b) Invert sugar
(c) Glycosidic linkage

3 × 1

26. For the first order thermal decomposition reaction, following data was obtained :



S. No.	Time(s)	Total Pressure (atm)
1	0	0.30
2	30	0.50

Calculate rate constant.

[Given, $\log 3 = 0.48$]

27. (a) Answer the following :

3 × 1

- (i) Why is the Equilibrium Constant (K_p) related to E°_{cell} and not to E_{cell} ?
(ii) Two metals 'A' and 'B' have standard electrode potential values of -0.24 V and $+0.80$ V respectively. Which of these will liberate hydrogen gas from dil. H_2SO_4 ?
(iii) Write the cell reaction which occurs in lead storage battery when it is in charging.

OR

- (b) What type of battery is Mercury cell ? Why it is more advantageous than dry cell ? Write overall reaction taking place in Mercury cell.

3

28. Calculate the boiling point of a solution containing 0.61 g of benzoic acid (Molar mass = 122 g mol^{-1}) in 5 g of CS_2 in which it dimerises to the extent of 88% . The boiling point and K_b of CS_2 are 46.2°C and $2.3 \text{ K kg mol}^{-1}$ respectively.

^56/1/2^



15

3) 3638 722379
28
[P.T.O.]

SECTION - D

The following questions are case-based questions. Each question has an internal choice and carries 4(2 + 1 + 1) marks each. Read the passage carefully and answer the questions that follow.

29. The Valence Bond Theory (VBT) explains the formation, magnetic behaviour and geometry of coordination compounds. The Crystal Field Theory (CFT) of coordination compounds is based on the effect of different crystal fields (provided by the ligands taken as point charges), on the degeneracy of d-orbital energies of the central metal atom/ion. The splitting of the d-orbitals provides different electronic arrangements in strong and weak crystal fields.

Answer the following questions :

- (a) In octahedral crystal field, energies of which d-orbitals will be raised when ligands approach the central metal atom/ion ? Give reason in support of your answer. 2
- (b) Using crystal field theory, write the electronic configuration of central metal atom/ion of the following : 1
- (i) $[\text{CoF}_6]^{3-}$
- (ii) $[\text{Co}(\text{NH}_3)_6]^{3+}$ [At. No. : Co = 27]
- (c) $[\text{NiCl}_4]^{2-}$ is paramagnetic while $[\text{Ni}(\text{CO})_4]$ is diamagnetic though both are tetrahedral. Why ? [Atomic No. : Ni = 28] 1

OR

- (c) Write hybridization and magnetic behaviour of the complex $[\text{Fe}(\text{CN})_6]^{3-}$. [Atomic No. : Fe = 26] 1

30. The reaction of amines with mineral acids to form ammonium salts shows that these are basic in nature. Aliphatic amines are stronger bases than ammonia whereas aromatic amines are weaker bases than ammonia. Aliphatic and aromatic primary and secondary amines react with acid chlorides, anhydrides and esters by nucleophilic substitution reaction. The main problem encountered during electrophilic substitution reactions of aromatic amines is that of their high reactivity. Substitution tends to occur at ortho- and para-positions. Hinsberg reagent is used for the identification and distinction between primary, secondary and tertiary amines. Aryldiazonium salts, usually obtained from arylamines, undergo replacement of the diazonium group with a variety of nucleophiles to provide advantageous methods for producing aryl halides, cyanides, phenols and arenes.

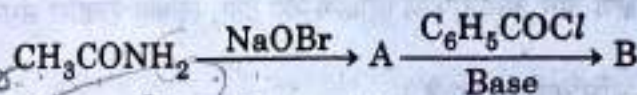


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Answer the following questions :

- (a) (i) Why $\text{CH}_3 - \text{NH}_2$ is a stronger base than $(\text{CH}_3)_3\text{N}$ in aqueous solution ?

- (ii) Write structural formulae of the compound A and B :



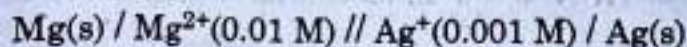
- (b) A compound 'X' with molecular formula $\text{C}_3\text{H}_9\text{N}$ reacts with Hinsberg reagent to give a product insoluble in alkali. Identify 'X'.

OR

- (b) How can you convert aniline to benzonitrile ?
(c) Why is $(-\text{NH}_2)$ group of aniline acetylated before carrying out nitration ?

SECTION - E

31. (a) Calculate emf and ΔG for the following cell at 298 K :

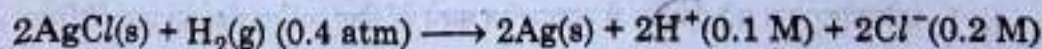


Given : $E^\circ_{\text{Mg}^{2+}/\text{Mg}} = -2.37 \text{ V}$ $E^\circ_{\text{Ag}^+/\text{Ag}} = +0.80 \text{ V}$

[$1 \text{ F} = 96500 \text{ C mol}^{-1}$, $\log 10 = 1$]

OR

- (b) For the reaction :



Calculate emf of the cell at 25 °C.

Given : $\Delta G^\circ = -43500 \text{ J mol}^{-1}$

[$\log 10 = 1$, $1 \text{ F} = 96500 \text{ C mol}^{-1}$]



$$\begin{array}{r} 0.80 \\ + 2.37 \\ \hline 3.17 \end{array}$$

[P.T.O.]

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32. (a) (i) An organic compound (X) has the molecular formula $C_6H_{10}O$.

Draw structures for (X) if it :

3 + 2

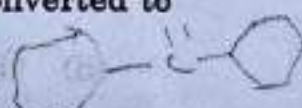
(I) does not give Tollen's test but gives a positive iodoform test. *carboxylic acid* *Aldehyde*

(II) does not give Tollen's test and iodoform test but undergoes Aldol condensation. *NaOH* *CHI₃*

(III) undergoes Cannizzaro's reaction.

(ii) Show how each of the following compounds can be converted to benzoic acid :

(I) Acetophenone (II) Ethyl benzene



OR

(b) Answer the following questions :

5 × 1

(i) Draw structure of the 2, 4-dinitrophenyl hydrazone derivative of benzaldehyde.

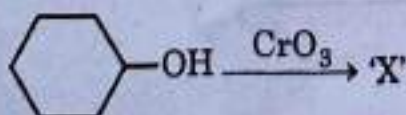
(ii) Arrange the following in increasing order of their reactivity towards HCN :

Di-tert. butyl ketone, Acetaldehyde, Acetone

(iii) Give a simple chemical test to distinguish between benzoic acid and ethyl benzoate.

(iv) Write the name of the reagent to convert Ethanenitrile to Ethanal.

(v) Draw the structure of 'X' in the following reaction :



33.

- (a) (i) From the given data of E° values, answer the following questions :

3 + 2

$E^\circ_{M^{2+}/M}$	V	Cr	Mn	Fe	Co	Ni	Cu
	-1.18	-0.91	-1.18	-0.44	-0.28	-0.25	+0.34

(I) Why $E^\circ_{M^{2+}/M}$ show irregular trend in the above values ?

(II) Why is $E^\circ_{Cu^{2+}/Cu}$ value exceptionally positive ?

(III) Why $E^\circ_{Mn^{2+}/Mn}$ value is highly negative ?

- (ii) Write the ionic equations for the oxidising action of potassium permanganate for its reaction with I^- in both acidic and alkaline solutions.

OR

- (b) Answer the following questions :

(i) Name a member of the lanthanoid series

(I) which exhibits +4 oxidation state

(II) which exhibits +2 oxidation state.

(ii) Why transition metals act as good catalyst ?

(iii) Why Cr has higher melting point than Mn ?

(iv) What happens when acidic solution of potassium permanganate is allowed to stand for sometime ? Give the equation involved. What is this type of reaction called ?

